CHEM- 102-002
General CHEM Tutorial II

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1 Credit Hr.
Tutorial: R 1.00 PM – 1.50 PM STC 135

Text/ Manual and other required material:

- **Text**: General Chemistry, 9 th Edition, Houghton Mifflin Company, by Ebbing.
- **Supplies:** Non-programmable Calculator (bring to class)

Course Description: The course will be cover and act as a support to understand the fundamental chemistry topics including chemical reaction rates, chemical equilibrium, acid base chemistry, solubility, thermodynamics, electrochemistry, nuclear chemistry, organic chemistry, inorganic chemistry and biochemistry.

Course Outcome: Use LeChatelier's principle to predict the effects of concentration, pressure and temperature changes on equilibrium mixtures. Balance Oxidation-Reduction reaction. Write simple structures of the organic compound and name the compound.

Class Procedure: The course is mainly to support lecture and problem solving session. Students are encouraged to do the problems at home and also in class and to work with me on doing the problems. Chemistry is a physical science and it is imperative to master calculations to pass the course.

Communication: <u>If the faculty needs to contact an individual student, it will be via the student's e-mail account</u>. Students should check e-mail frequently, especially after absence. E-mail is the best, easiest and fastest way to communicate with me since I check my email daily.

Course Requirements:

1. ***Exams*:** Three unit exams and one mandatory, comprehensive final exam. If you missed an exam, the points for the missed exam will be replaced by final exam grade making the final exam count for 15 percentage of your grade. **No make-up exams**. All students must take the exams at the scheduled time. Exams cover lecture, quiz and homework.
2. ***Quizzes*:** There will be quizzes, which will be 20 % of the overall class grade. Each week we might have announced or unannounced quizzes, which will be given at the class. Lowest two quiz scores will be dropped. There will be no makeup quizzes. Quizzes cover lecture and homework material.
3. ***Attendance and Class Participation:*** Class attendance is required to pass the course, so you should attend each class. The Department of Chemistry adheres to the attendance policy set by the University as stated in the most current Undergraduate Catalog. The attendance record is taken from the daily sign-in sheet will be counted as a missing lecture. Excessive absence is defined as missing more than 10% of the lectures or more than 10% of the laboratory sessions without excusable reasons. Excessive absence will be reported to the Dean of the College and Dean of students. In addition, according to the TAMU-Commerce Procedure A13.02, if a student

has excessive absences, the instructor may drop the student from the course. The instructor will only excuse an absence if the student provides, with appropriate document an excusable reason allowed by the TAMU-Commerce Procedure A13.02. Good class attendance will be necessary in order to pass this course.

Methods of Student Evaluation and Grading Scale: 3 exams, one final exam, quizzes and homework will evaluate Students. Three Exams (each 15 %) will be 45%, Final Exam 20 %, quizzes 20% and homework 15 %. The grade is based on a weighted average.

The grade scale will be A= (86.0 - 100%), B = (73.0 – 85.9 %), C = (60-72.9%), D = (45-59.9%), F = <45%.

Class: Students are required to turn off all cell phones, MP3 players, PDA's, Pagers, and any other electronic devices before entering the class or in the laboratory. Students are expected to comply with the student code of conduct as stated Student's Guidebook, Policies and Procedures, Conduct. If the student's failed to comply with the code of conduct and being disrespectful, disruptive to the instructor or the students of the class, the instructor reserves the right to dismiss the student from the class on the first offense. A second offense may constitute dismissal from the course with a failing grade.

Academic Integrity and Dishonesty Policy: All students are expected to pursue their scholastic careers with honesty and integrity. Academic dishonesty includes (but is not limited to) cheating, falsification of date, plagiarism, and contracting/collusion with others to take your tests or do your work. Cheating is the use or acquisition of information (data, constants, formulas, textual material, etc.) from either unauthorized sources or in an unauthorized manner.

Examples include but are not limited to: -

- Exchanging information during a test or quiz
- Looking at another student's paper during a test or quiz
- Bringing information in any forms into a test or quizzes other than personal knowledge. This includes written notes (crib sheets) and digitally stored information (formulas, constants, textual, etc.) on calculators, cell phones, pagers etc.
- Looking at a book or other unauthorized source during the test or quiz.
- Accessing information by any electronic means (cellular phone, pagers, personal stereos, etc.)
- Processing data or information in an unauthorized manner using a programmable calculator or computer. In other words, unless you have received authorization, you are not to use any computer program. This includes specialty computers or calculators in which the programming is built in to the computer; you are permitted to use simple calculators, which perform arithmetical, Logarithmic, and trigonometric functions.

Disciplinary action will be pursued in all instances in which it is determined that academic dishonesty has occurred. Disciplinary action may include but is not limited to:

- Assignment of a failing grade for a test, examination, or assignment;
- Assignment of a failing grade for a course;
- Student disciplinary sanction.

Student Withdrawal: It is the student's responsibility to withdraw from class if so desired. However, the instructor reserves the right to administratively withdraw any student who is not actively fulfilling the objectives of the course before the final.

Incomplete: An incomplete is given only when a student, for a valid reason, has been unable to complete course within the time allotted and has a current average ≥ 70 %. This is not allowed except in documented illness.

Tentative Tutorial Schedule
CHEM-102-002; Fall 2012

Week	Date	Thursday Tutorial
1	**8/30/12**	Class Syllabus and Policy Ch. 12: Solutions
2	**9/6/12**	Ch. 12: Solutions
3	**9/13/12**	Ch. 13: Chemical Equilibrium Ch. 14: Rates of Reaction **Exam I (Chapter 12-14)**
4	**9/20/12**	Ch. 14: Rates of Reaction
5	**9/27/12**	Ch. 15: Acids and Bases Ch. 16: Acid-Base Equilibrium
6	**10/4/12**	Ch. 16: Acid-Base Equilibrium Ch. 17: Solubility and Complex-Ion Equilibrium **Exam II (Chapter 15-17)**
7	**10/11/12**	Ch. 17: Solubility and Complex-Ion Equilibrium
8	**10/18/12**	Ch. 18: Thermodynamics and Equilibrium
9	**10/25/12**	Ch. 18: Thermodynamics and Equilibrium Ch. 19: Electrochemistry
10	**11/1/12**	**Exam III (Chapter 18-19)**
11	**11/8/12**	Ch. 19: Electrochemistry
12	**11/15/12**	Ch. 23: Organic Chemistry
13	**11/22/12**	Ch. 23: Organic Chemistry Ch. 20: Nuclear Chemistry
14	**11/29/12**	Ch. 21-22 Main Group and Transition Elements
15	**12/6/12**	**Final Exam (Chapter 12-23)**

Note: Instructor keeps the right to make any changes of the syllabus.

Important dates:

1. Exam I– 13th September, 2012
2. Exam II– 4th October, 2012
3. Exam III – 1st November, 2012
4. Final Exam – 6th December, 2012