

CHEM- 101-001
General CHEM Tutorial I

Faculty: Dr. Tasneem Hossain-Kumar

1 Credit Hr.

Office Location: STC 302

Tutorial: R 1.00 PM – 1.50 PM STC 127

Office Phone: 903-886-5399

Office Hours: T 3.20- 4.50 PM and by appointment

Email: tasneemhossain@gmail.com

Tasneem.Hossain-Kumar@tamuc.edu

Text/ Manual and other required material:

- **Text:** General Chemistry, 9 th Edition, Houghton Mifflin Company, by Ebbing.
- **Supplies:** Non-programmable Calculator (bring to class)

Course Description: The course will be cover and act as a support to understand the fundamental chemistry topics including chemical reaction rates, chemical equilibrium, acid base chemistry, solubility, thermodynamics, electrochemistry, nuclear chemistry, organic chemistry, inorganic chemistry and biochemistry.

Course Outcome: Students will be able to report the answer with correct no of significant figure after performing mathematical calculation. They will be able to write the formula of the given compound. They will also be able to write the Lewis formula or electron dot structure of the given compound and predict the geometry of the molecule by VSEPR method.

Class Procedure: The course is mainly to support lecture and problem solving session. Students are encouraged to do the problems at home and also in class and to work with me on doing the problems. Chemistry is a physical science it is imperative to master calculations to pass the course.

Communication: If the faculty needs to contact an individual student, it will be via the student's e-mail account. Students should check e-mail frequently, especially after absence. E-mail is the best, easiest and fastest way to communicate with me since I check my email daily.

Course Requirements:

1. **Exams:** One unit exam and one mandatory, comprehensive final exam. If you missed **an exam**, the points for the missed exam will be replaced by final exam grade. **No make-up exams.** All students must take the exams at the scheduled time. Exams cover lecture, quiz and homework.
2. **Attendance and Class Participation:** Class attendance and class work is required to pass the course because it is 50% part of your final grade, so you should attend each class and participate in class work. The class work will be graded and will be signed by the instructor at the end of the class. The Department of Chemistry adheres to the attendance policy set by the University as stated in the most current Undergraduate Catalog. The attendance record is taken from the daily sign-in sheet will be counted as a missing lecture. Excessive absence is defined as missing more than 10% of the lectures or more than 10% of the laboratory sessions without excusable reasons. Excessive absence will be reported to the Dean of the College and Dean of students. In addition, according to the TAMU-Commerce Procedure A13.02, if a student has excessive absences, the instructor may drop the student from the course. The instructor will only excuse an absence if the

student provides, with appropriate document an excusable reason allowed by the TAMU-Commerce Procedure A13.02. Good class attendance will be necessary in order to pass this course.

Methods of Student Evaluation and Grading Scale: One exam (30 min), one final exam (50 min), and attendance and class work will evaluate Students. One Exam (25 %), Final Exam (25 %) and attendance and class work (50 %) will be of the final grade. The grade is based on a weighted average.

The grade scale will be A= (89.0 - 100%), B = (73.0 – 88.9 %), C = (60-72.9%), D = (45-59.9%), F = <45%.

Class: Students are required to turn off all cell phones, MP3 players, PDA's, Pagers, and any other electronic devices before entering the class or in the laboratory. Students are expected to comply with the student code of conduct as stated Student's Guidebook, Policies and Procedures, Conduct. If the student's failed to comply with the code of conduct and being disrespectful, disruptive to the instructor or the students of the class, the instructor reserves the right to dismiss the student from the class on the first offense. A second offense may constitute dismissal from the course with a failing grade.

Academic Integrity and Dishonesty Policy: All students are expected to pursue their scholastic careers with honesty and integrity. Academic dishonesty includes (but is not limited to) cheating, falsification of date, plagiarism, and contracting/collusion with others to take your tests or do your work. Cheating is the use or acquisition of information (data, constants, formulas, textual material, etc.) from either unauthorized sources or in an unauthorized manner.

Examples include but are not limited to: -

- Exchanging information during a test or quiz
- Looking at another student's paper during a test or quiz
- Bringing information in any forms into a test or quizzes other than personal knowledge. This includes written notes (crib sheets) and digitally stored information (formulas, constants, textual, etc.) on calculators, cell phones, pagers etc.
- Looking at a book or other unauthorized source during the test or quiz.
- Accessing information by any electronic means (cellular phone, pagers, personal stereos, etc.)
- Processing data or information in an unauthorized manner using a programmable calculator or computer. In other words, unless you have received authorization, you are not to use any computer program. This includes specialty computers or calculators in which the programming is built in to the computer; you are permitted to use simple calculators, which perform arithmetical, Logarithmic, and trigonometric functions.

Disciplinary action will be pursued in all instances in which it is determined that academic dishonesty has occurred. Disciplinary action may include but is not limited to:

- Assignment of a failing grade for a test, examination, or assignment;
- Assignment of a failing grade for a course;
- Student disciplinary sanction.

Student Withdrawal: It is the student's responsibility to withdraw from class if so desired. However, the instructor reserves the right to administratively withdraw any student who is not actively fulfilling the objectives of the course before the final.

**Tentative Tutorial Schedule
CHEM-101-001; Spring 2013**

Week	Date	Thursday Tutorial
1	1/17/13	Class Syllabus and Policy Ch. 1: Chemistry and Measurement
2	1/24/13	Ch. 1: Chemistry and Measurement Ch. 2: Atoms, Molecules and Ions
3	1/31/13	Ch. 2: Atoms, Molecules and Ions Ch. 3: Calculations with Chemical Formulas and Equations Mid Term Exam
4	2/7/13	Ch. 3: Calculations with Chemical Formulas and Equations Ch. 4: Chemical Reactions
5	2/14/13	Ch. 4: Chemical Reactions Ch. 5: The Gaseous State
6	2/21/13	Ch. 5: The Gaseous State Ch. 6: Thermochemistry
7	2/28/13	Ch. 6: Thermochemistry Ch. 7: Quantum Theory of the Atom
8	3/7/13	Ch. 7: Quantum Theory of the Atom Ch. 8: Electron Configurations and Periodicity
9	3/14/13	Ch. 8: Electron Configurations and Periodicity Ch. 9: Ionic and Covalent Bonding
10	3/21/13	Ch. 9: Ionic and Covalent Bonding
11	3/28/13	Ch. 9: Ionic and Covalent Bonding Ch.10: Molecular Geometry and Chemical Bonding Theory
12	4/11/13	Ch.10: Molecular Geometry and Chemical Bonding Theory Ch. 11: States of Matter; Liquids and Solids
13	4/18/13	Ch. 11: States of Matter; Liquids and Solids
14	4/25/13	Ch. 11: States of Matter; Liquids and Solids Review
15	5/2/13	Final Exam

Note: Instructor keeps the right to make any changes of the syllabus.

Important dates:

1. Exam I– 31st January, 2013
2. **Final Exam – 2nd May, 2013**